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## Equilibrium Continued - Solutions

Equilibrium in chemical reactions is dynamic because the forward and reverse reactions are occurring continuously and simultaneously at the same rates. Placing a stress on any equilibrium system, whether it is chemical, biological, societal, environmental, or personal, causes the equilibrium position to change.

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Mrs. Zuberbuehler - Mrs. Zuberbuehler

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POGIL™ Activities for High School Chemistry. There is a downside to checkpoints. Most students are expecting the equilibrium to be when they have the same amount of water in both cylinders.

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- Chemical equilibrium occurs when opposing reactions are proceeding at equal rates.
- The rate at which the products are formed from the reactants equals the rate at which the reactants are formed from the products.

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## CHEMICAL EQUILIBRIUM (ICE METHOD)

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4™ Activities for AP\* Chemistry POGIL 12. Do all buffers keep solutions at a neutral pH? Justify your answer with data from Model 2.  
13. Calculate the pK a 's for the weak acids in each of the buffer

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solutions described in Model 2 and list them in the model. 14. How are the  $pK_a$  values of the weak acids related to the pH of the buffer ...

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Buffers - SUPERTALLTEACHER

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AHS Chemistry Resource Site - Unit 6 - Rates & Equilibrium

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Dynamic Equilibrium Pogil Answers at equilibrium:  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + \text{energy (heat)}$  In Table 1 apply Le Chatelier's Principle and indicate the direction of the shift in equilibrium if the indicated stress is applied to the reaction system.

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reaction quotient. The key to knowing which direction a reaction will  
need to proceed in order to reach equilibrium is knowing if you have  
too much reactant or too much The reaction quotient for a reaction is



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the ratio of products to reactants, similar to the equilibrium constant.

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